

# Simulations of ionic liquids confined by metal electrodes using periodic Green functions

Matheus Giroto, Alexandre P. dos Santos, and Yan Levin

Citation: *The Journal of Chemical Physics* **147**, 074109 (2017); doi: 10.1063/1.4989388

View online: <https://doi.org/10.1063/1.4989388>

View Table of Contents: <http://aip.scitation.org/toc/jcp/147/7>

Published by the [American Institute of Physics](#)

---

## Articles you may be interested in

[Simulations of Coulomb systems confined by polarizable surfaces using periodic Green functions](#)

*The Journal of Chemical Physics* **147**, 184105 (2017); 10.1063/1.4997420

[Generalized Einstein relation for the mutual diffusion coefficient of a binary fluid mixture](#)

*The Journal of Chemical Physics* **147**, 074902 (2017); 10.1063/1.4993958

[Perspective: Surface freezing in water: A nexus of experiments and simulations](#)

*The Journal of Chemical Physics* **147**, 060901 (2017); 10.1063/1.4985879

[Lattice model of ionic liquid confined by metal electrodes](#)

*The Journal of Chemical Physics* **148**, 193829 (2018); 10.1063/1.5013337

[Electrostatic and induction effects in the solubility of water in alkanes](#)

*The Journal of Chemical Physics* **147**, 074506 (2017); 10.1063/1.4997916

[Thermodynamics of atomistic and coarse-grained models of water on nonpolar surfaces](#)

*The Journal of Chemical Physics* **147**, 074702 (2017); 10.1063/1.4999337

---

PHYSICS TODAY

WHITEPAPERS

### ADVANCED LIGHT CURE ADHESIVES

Take a closer look at what these environmentally friendly adhesive systems can do

READ NOW

PRESENTED BY  
 MASTERBOND  
ADHESIVES | SEALANTS | COATINGS

# Simulations of ionic liquids confined by metal electrodes using periodic Green functions

Matheus Giroto,<sup>a)</sup> Alexandre P. dos Santos,<sup>b)</sup> and Yan Levin<sup>c)</sup>

*Instituto de Física, Universidade Federal do Rio Grande do Sul, Caixa Postal 15051, CEP 91501-970 Porto Alegre, RS, Brazil*

(Received 8 June 2017; accepted 1 August 2017; published online 17 August 2017)

We present an efficient method for simulating Coulomb systems confined by metal electrodes. The approach relies on Green function techniques to obtain the electrostatic potential for an infinite periodically replicated system. This avoids the use of image charges or an explicit calculation of the induced surface charge, both of which dramatically slows down the simulations. To demonstrate the utility of the new method, we use it to obtain the ionic density profiles and the differential capacitances, which are of great practical and theoretical interest, for a lattice model of an ionic liquid. *Published by AIP Publishing.* [<http://dx.doi.org/10.1063/1.4989388>]

## I. INTRODUCTION

Simulations of Coulomb systems in confined geometries with a reduced symmetry are notoriously difficult. This is due to the long range nature of the Coulomb interaction, which prevents the use of simple periodic boundary conditions. Instead, an infinite number of replicas must be considered so that each particle in the simulation cell interacts with an infinite set of periodic replicas of itself and of all the other ions. To efficiently sum over the replicas, the usual approach relies on Ewald summation techniques.<sup>1–5</sup> Ewald methods have been implemented for both Coulomb and gravitational systems in 3-d, and various optimization techniques have been developed. Unfortunately, when the symmetry of the system is reduced, which is the case when an interface is present, the computational cost of summation over replicas increases dramatically due to the appearance of special functions and slow convergence.<sup>6–8</sup> To overcome these problems, a number of approaches have been proposed.<sup>9–15</sup> The difficulty is that these methods are not easily generalized to systems bounded by metallic or dielectric surfaces. The dielectric interfaces are important in many biophysics applications, while the metallic electrodes are omnipresent in electrochemistry and play a fundamental role in the discussion of ionic liquids which, due to their use in renewable energy storage devices, are of great practical and technological importance.<sup>16–19</sup> To simulate metallic surfaces of electrodes, various approaches have been proposed.<sup>20–22</sup> Unfortunately, all are computationally very expensive, relying on a minimization procedure to calculate the induced surface charge at every molecular dynamics time step.<sup>23,24</sup> In the present paper, we propose a completely new method for studying ionic systems confined by parallel metal surfaces. The formalism is based on the periodic Green functions which are constructed to satisfy the appropriate Dirichlet boundary conditions. The resulting infinite sum over replicas is fast

convergent and can be truncated at a reasonably small number of replicas. To show the utility of the method, we apply it to calculate the ionic density profiles and the differential capacitances of a lattice model of a room temperature ionic liquid, confined by two metal electrodes.<sup>24–28</sup>

## II. THE THEORY

Our goal is to calculate the electrostatic potential inside a simulation cell of size  $L_x \times L_y \times L$ , bounded by conducting surfaces separated by distance  $L$ . For an ion located at  $(x_0, y_0, z_0)$  inside the main simulation cell, there will be an infinite set of replicas located at  $(x_0 \pm m_x L_x, y_0 \pm m_y L_y, z_0)$ , with  $m_x, m_y \in \mathbb{Z}^+$ , see Fig. 1.

We start by calculating the electrostatic potential at position  $\mathbf{r} = \rho\hat{\rho} + \phi\hat{\phi} + z\hat{z}$  produced by a *single* ion of charge  $q$  located at  $\mathbf{r}' = z_0\hat{z}$ , see Fig. 1, between two parallel grounded infinite metal electrodes. To explore the azimuthal symmetry of the potential, we use a cylindrical coordinates  $(\rho, \phi, z)$ . To obtain the electrostatic potential requires us to solve the Poisson equation:<sup>29</sup>

$$\nabla^2 \phi(\mathbf{r}, \mathbf{r}') = -\frac{4\pi q}{\epsilon} \delta(\mathbf{r} - \mathbf{r}'), \quad (1)$$

with the Dirichlet boundary condition  $\psi_0 = 0$  at each surface. We start by expanding the delta function in the eigenfunctions of the differential operator

$$\frac{d^2 \psi_n}{dz^2} + k_n^2 \psi_n = 0, \quad (2)$$

satisfying the boundary conditions  $\psi_n(0) = \psi_n(L) = 0$ . The eigenfunctions are found to be  $\psi_n(z) = \sqrt{2/L} \sin(k_n z)$ , with  $k_n = n\pi/L$ . The delta function can then be written as

$$\delta(z - z_0) = \frac{2}{L} \sum_{n=1}^{\infty} \sin\left(\frac{n\pi z}{L}\right) \sin\left(\frac{n\pi z_0}{L}\right). \quad (3)$$

The electrostatic potential can now be written as

<sup>a)</sup>Electronic mail: matheus.giroto@ufrgs.br

<sup>b)</sup>Electronic mail: alexandre.pereira@ufrgs.br

<sup>c)</sup>Electronic mail: levin@if.ufrgs.br

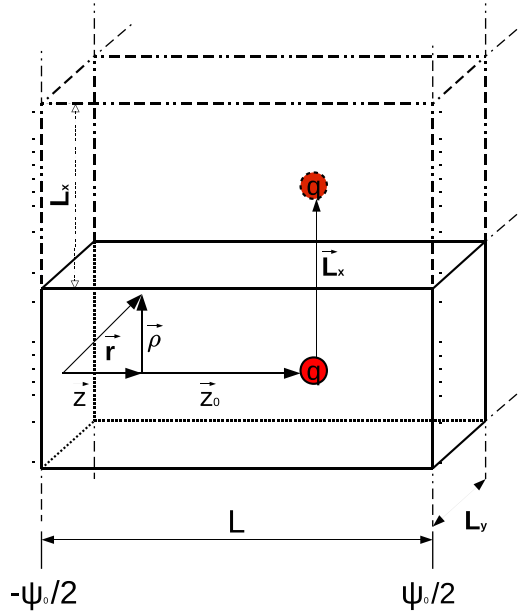


FIG. 1. Point charge  $q$  located at  $z_0\hat{z}$  between two infinite metal electrodes with constant potential difference  $\psi_0$ . The electrodes are at  $z = 0$  and  $z = L$ . The electrostatic potential is calculated at a point  $\mathbf{r}$  indicated in cylindrical coordinates. The dashed lines show the first replica of the simulation box in the  $x$  direction.

$$\phi(\rho, z; z_0) = \frac{2q}{\epsilon L} \sum_{n=1}^{\infty} \sin\left(\frac{n\pi z}{L}\right) \sin\left(\frac{n\pi z_0}{L}\right) g_n(\rho). \quad (4)$$

Substituting this expression into Eq. (1), we obtain an ordinary differential equation for  $g_n(\rho)$ ,

$$\frac{1}{\rho} \frac{d}{d\rho} \left( \rho \frac{dg_n}{d\rho} \right) - k_n^2 g_n = -\frac{2}{\rho} \delta(\rho), \quad (5)$$

which has modified Bessel functions of order zero as solutions,  $g_n(\rho) = AI_0(k_n\rho) + BK_0(k_n\rho)$ . Since the potential must vanish as  $\rho \rightarrow \infty$ , the coefficient  $A = 0$ , while the coefficient  $B$  is determined by the singular part of the potential and is found to be  $B = 2$ . The electrostatic potential produced by an ion located at  $\mathbf{r} = z_0\hat{z}$  between two grounded metal surfaces is then<sup>29</sup>

$$\phi(\rho, z; z_0) = \frac{4q}{\epsilon L} \sum_{n=1}^{\infty} \sin(k_n z) \sin(k_n z_0) K_0(k_n \rho). \quad (6)$$

For simulating an ionic system, we will need to periodically replicate the main simulation cell. Since the electrostatic potential in Eq. (6) satisfies the appropriate boundary conditions, the electrostatic potential produced by an ion located at  $(x_0, y_0, z_0)$  and all of its periodic replicas can be simply obtained by a superposition. We find

$$G(\mathbf{r}; \mathbf{r}_0) = \frac{4q}{\epsilon L} \sum_{m=-\infty}^{\infty} \sum_{n=1}^{\infty} \sin(k_n z) \sin(k_n z_0) \times K_0 \left( k_n \sqrt{(x - x_0 + m_x L_x)^2 + (y - y_0 + m_y L_y)^2} \right). \quad (7)$$

The modified Bessel function  $K_0(x)$  decays exponentially for large  $x$ ; therefore, in practice, we will need only a small number of replicas to obtain the electrostatic potential to any desired

accuracy. Unfortunately, Eq. (7) is ill defined when  $x = x_0$  and  $y = y_0$ . The problem arises because  $K_0(x)$  has a logarithmic divergence at  $x = 0$ , which manifests itself in  $m_x = m_y = 0$  term of Eq. (7). This term corresponds to the electrostatic potential arising from the ion inside the main cell. As long as  $z \neq z_0$ , the electrostatic potential will remain finite as  $x \rightarrow x_0$  and  $y \rightarrow y_0$ , this limit, however, is very difficult to obtain because of a slow convergence of the sum in Eq. (6). To overcome this difficulty, we will use a different representation of the Green function to calculate the electrostatic potential produced by the ion inside the main simulation cell.

Once again, we consider *one* (no replicas) ion located between two infinite grounded conducting surfaces at  $\mathbf{r} = z_0\hat{z}$ . We now use the following representation of the delta function:<sup>30</sup>

$$\frac{1}{\rho} \delta(\rho) = \int_0^{\infty} k J_0(k\rho) dk, \quad (8)$$

where  $J_0$  is the Bessel function of order zero. The electrostatic potential can now be written as

$$\phi(\rho, z; z_0) = \frac{q}{\epsilon} \int_0^{\infty} k J_0(k\rho) g_k(z, z_0) dk. \quad (9)$$

Substituting Eq. (9) into Eq. (1), we obtain an ordinary differential equation for  $g_k(z, z_0)$ :

$$\frac{d^2 g_k}{dz^2} - k^2 g_k = -2\delta(z - z_0). \quad (10)$$

Applying the boundary conditions, we finally obtain

$$\phi(\rho, z; z_0) = \frac{q}{\epsilon} \int_0^{\infty} dk J_0(k\rho) \times \frac{e^{k|z-z_0|-2kL} + e^{-k|z-z_0|} - e^{-k(z+z_0)} - e^{k(z+z_0)-2kL}}{1 - e^{-2kL}}, \quad (11)$$

which is well behaved when  $\rho \rightarrow 0$ , as long as  $z \neq z_0$ . This expression is equivalent to Eq. (4) and can be used to calculate the electrostatic potential produced by the ion inside the main simulation cell, replacing the  $m_x = m_y = 0$  term of Eq. (7), since the integral in Eq. (11) will rapidly converge even if the potential is calculated at  $\rho = 0$ .

Suppose an ion is placed at  $z = z_0$  between two infinite grounded metal surfaces. How much charge will be induced on each electrode? The surface charge density on the left electrode is

$$\sigma(\rho) = -\frac{\epsilon}{4\pi} \frac{\partial \phi}{\partial z} \Big|_{z=0} = -\frac{q}{2\pi} \int_0^{\infty} dk J_0(k\rho) k \frac{\sinh[k(L - z_0)]}{\sinh(kL)}, \quad (12)$$

and the total charge is

$$Q_l^0 = -\frac{q}{2\pi} \int d\varphi \int \rho d\rho \int dk k J_0(k\rho) \frac{\sinh[k(L - z_0)]}{\sinh(kL)}. \quad (13)$$

Equation (13) is conditionally convergent. To conveniently perform the integral, we introduce a convergence factor  $e^{-\alpha\rho}$ , which allows us to change the order of integration. Performing the integration first over  $\rho$  and then changing variables and taking the  $\alpha \rightarrow 0$  limit, we find that the total charge on the left electrode is

$$Q_l^0 = -q \left( 1 - \frac{z_0}{L} \right). \quad (14)$$

Similarly the surface charge on the right electrode is  $Q_r^0 = -qz_0/L$ .

So far, our discussion has been restricted to the grounded metal surfaces. Often, however, the electrostatic potential difference between the electrodes is controlled by an external battery so that the potential of the electrode located at  $z = 0$  is fixed at  $-\psi_0/2$  and of the electrode located at  $z = L$  at  $+\psi_0/2$ . Using the uniqueness property of the Laplace equation, it is simple to account for the surface potential controlled by an external source. We observe that if we add to Eq. (7) a potential

$$\phi_s(z) = \left(\frac{z}{L} - \frac{1}{2}\right)\psi_0, \quad (15)$$

the sum will satisfy the Laplace equation with inhomogeneous boundary conditions, providing a unique solution to the problem. For a periodically replicated charge neutral system with  $N$  ions at positions  $\{\mathbf{r}_i\}$  and electrodes held at potentials  $\mp\psi_0/2$ , respectively, the total charge on the left and right electrodes *within the simulation cell* will then be

$$Q_{l,r} = \mp \frac{\epsilon\psi_0 A}{4\pi L} \pm \sum_{i=1}^N q_i \frac{z_i}{L}, \quad (16)$$

where  $A = L_x L_y$  is the area of the electrode inside the simulation cell. Note that  $Q_l = -Q_r$ .

### III. SIMULATIONS

We are now in a position to perform simulations of  $N$ -body Coulomb systems confined by two parallel metal electrodes. The object of particular interest for the ionic liquids community is the differential capacitance, which can be obtained from the fluctuations of the surface charge on the electrodes.<sup>20</sup> The partition function in the fixed electrostatic potential ensemble is

$$\mathcal{Z}_\psi = \int \prod_{i=1}^N d\mathbf{r}_i \int dQ e^{-\beta[E(\mathbf{r}_1, \dots, \mathbf{r}_N, Q) - \psi Q]}, \quad (17)$$

where  $\beta = 1/k_B T$  and the surface charge on the left and right electrodes is  $\mp Q$ , respectively. Note that in this ensemble, the surface charge on the electrodes is allowed to fluctuate. The differential capacitance of the system can then be calculated straightforwardly as

$$C = \frac{1}{A} \frac{\partial \langle Q \rangle}{\partial \psi} = \frac{1}{\beta A} \left( \frac{\partial^2 \ln \mathcal{Z}_\psi}{\partial \psi^2} \right) = \frac{\beta}{A} [\langle Q^2 \rangle - \langle Q \rangle^2]. \quad (18)$$

It is important to note that in order to perform a simulation at a *fixed electrostatic potential*, we need to know the total electrostatic energy  $E(Q)$  of a system with electrodes carrying a *fixed amount* of surface charge  $-Q$  and  $+Q$ , respectively. Since the electrodes are metallic, they must be equipotential. This means that the distribution of the surface charge will *not* be uniform and will respond to ionic motion. For a given  $Q$ , the surface potential  $\psi_0$  will, therefore, fluctuate. Since the system is charge neutral, the surface potential for a given ionic distribution inside the simulation cell can be easily calculated using Eq. (16),

$$\psi_0 = \frac{4\pi L}{\epsilon A} \left( Q + \sum_{i=1}^N q_i \frac{z_i}{L} \right). \quad (19)$$

The total electrostatic energy inside the simulation cell is then

$$E(Q) = \frac{1}{2} \sum_{i \neq j}^N q_i G(\mathbf{r}_i; \mathbf{r}_j) + \sum_{i=1}^N \left[ U_s(\mathbf{r}_i) + \frac{1}{2} q_i \phi_s(z_i) \right] + \frac{1}{2} \psi_0 Q, \quad (20)$$

where the periodic Green function is given by Eq. (7) with  $m_x = m_y = 0$  term replaced by Eq. (11), and the self energy of an ion at  $\mathbf{r}_i$  is

$$U_s(\mathbf{r}_i) = \frac{q_i}{2} \lim_{\rho \rightarrow 0} \left[ G(\mathbf{r}_i; \mathbf{r}_i) - \frac{q_i}{\epsilon \rho} \right]. \quad (21)$$

Using the identity

$$\int_0^\infty dk J_0(k\rho) = \frac{1}{\rho}, \quad (22)$$

the limit in Eq. (21) can be performed explicitly,<sup>30</sup> resulting in

$$U_s(\mathbf{r}_i) = \frac{q^2}{2\epsilon} \int dk \frac{2e^{-2kL} - e^{-2kz_i} - e^{2kz_i - 2kL}}{1 - e^{-2kL}} + \frac{2q^2}{\epsilon L} \sum_{m \neq 0} \sum_{n=1}^\infty \sin^2(k_n z_i) K_0 \left( k_n \sqrt{m_x^2 L_x^2 + m_y^2 L_y^2} \right). \quad (23)$$

In practice, since  $K_0(x)$  decays exponentially for large  $x$ , the sums in Eq. (23) converge very fast. To demonstrate the utility of the present method, we study a Coulomb lattice gas<sup>31–33</sup> confined between two electrodes held at potential  $\mp\psi/2$ , respectively. The Monte Carlo simulations are performed using the Metropolis algorithm. To further speed up the simulations, we have pre-calculated the electrostatic potentials at each lattice position at the beginning of the simulation. We allow swap moves between the ions and between the ions and the empty sites, during which the surface charge on the electrode remains fixed and the energy of the system  $E(Q)$  is calculated using Eq. (20). We also allow moves in which the surface charge on the electrodes increases or decreases in accordance with the Boltzmann factor of Eq. (17).

The simulations are performed in a cell of volume  $V = L_x L_y L$ , with  $L_x = L_y = 80 \text{ \AA}$  and  $L = 3L_x$ . The lattice gas is confined in the region  $-L_x/2 < x < L_x/2$ ,  $-L_y/2 < y < L_y/2$ , and  $0 < z < L$ . The negatively charged electrode is positioned at  $z=0$  and the positive one at  $z=L$ . We define the Bjerrum length as  $\lambda_B = q^2/k_B T \epsilon$  and consider two specific values  $\lambda_B = 7.2 \text{ \AA}$  and  $\lambda_B = 38.4 \text{ \AA}$ . The first value is appropriate for room temperature electrolytes while the second is for room temperature ionic liquids,<sup>34–36</sup> which have dielectric constant around  $\epsilon = 15$ . The concentration of ionic liquid is controlled by the capacity factor  $\gamma = (N_+ + N_-)/(N_+ + N_- + N_0)$ , where  $N_+$  is the number of cations,  $N_-$  is the number of anions, and  $N_0$  is the number of voids. We will set  $\gamma$  to  $\frac{1}{20}$  for electrolytes and  $\frac{1}{2}$  for ionic liquids. The lattice spacing is set to  $8 \text{ \AA}$ , a characteristic of ionic diameter. For now, we consider a symmetric case with charge of cation  $q$  and charge of anions  $-q$ , where  $q$  is the charge of the proton. The model, however, can be easily extended to asymmetric ionic liquids. In the simulations, we have used around  $\approx 10^4$

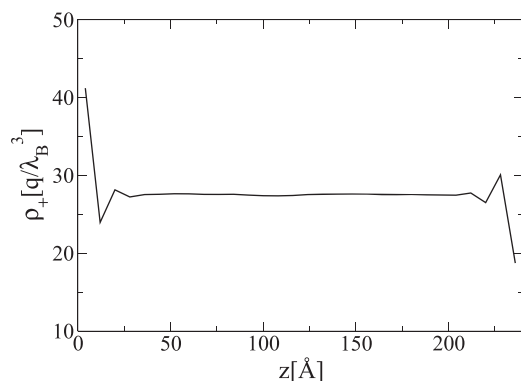


FIG. 2. Cationic profile of ionic liquid between the electrodes. The parameters are  $\gamma = \frac{1}{2}$ ,  $\lambda_B = 38.4 \text{ \AA}$ , and  $\sigma = 0.05 \text{ V}$ .

$m$ -vectors in the energy computation. The averages were calculated with  $5 \times 10^4$  uncorrelated samples after equilibrium was achieved.

In Fig. 2, we show the oscillatory behavior of the counterion density profile near an electrode,<sup>24</sup> and in Fig. 3, we present the differential capacitance in electrolyte and ionic liquid regimes. Figure 3(a) shows the characteristic minimum of differential capacitance at zero potential predicted by the Poisson-Boltzmann theory, followed by a maximum for higher applied voltages. The behavior is a characteristic of electrolyte solutions.<sup>37</sup> On the other hand, in the regime of ionic liquids, where steric and electrostatic correlations play the dominant

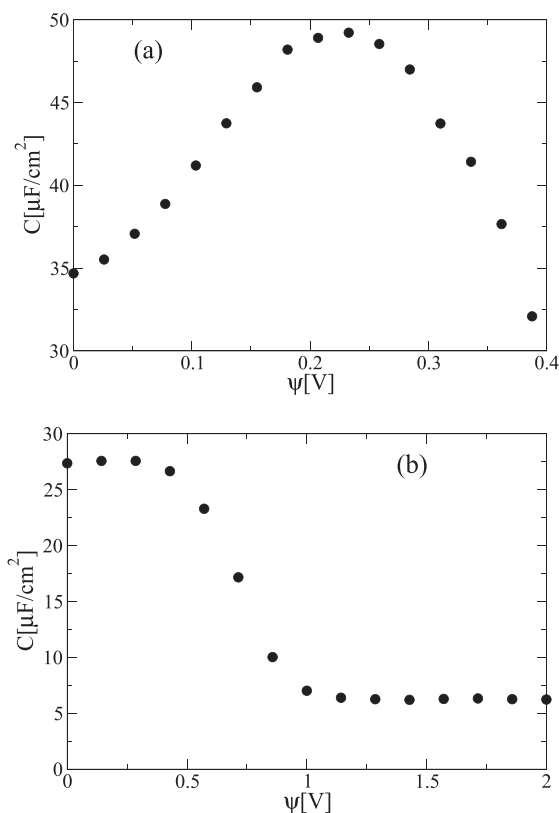


FIG. 3. Differential capacitance calculated using Eq. (18). Panel (a) shows the electrolyte regime with parameters  $\gamma = \frac{1}{20}$  and  $\lambda_B = 7.2 \text{ \AA}$ ; and panel (b) shows the typical bell-shaped differential capacitance of ionic liquids,  $\gamma = \frac{1}{2}$  and  $\lambda_B = 38.4 \text{ \AA}$ .

role,<sup>38</sup> the behavior is quite different.<sup>39–41</sup> Figure 3(b) shows that unlike electrolytes, ionic liquids have a maximum of differential capacitance at  $\psi = 0 \text{ V}$ . It is gratifying to see that a simple lattice model captures this complicated transition of differential capacitance between electrolyte and ionic liquid regimes.

#### IV. CONCLUSIONS

We have presented a new method for simulating ionic systems confined by infinite metal electrodes. Our algorithm is based on periodic Green functions derived in the present letter. The main advantage of the method is that it avoids the explicit calculation of the potential produced by the infinite distribution of the image charges or a numerical calculation of the induced surface charge at each simulation time step. Furthermore, since the potential produced by the ions is effectively screened by the electrodes, we only need a small number of replicas to achieve any desired precision. As a demonstration of the utility of the method, we have applied it to the calculations of the differential capacitance of a Coulomb gas, both in the electrolyte and ionic liquid regimes. In the future work, we will extend these calculations to continuum ionic liquids.

#### ACKNOWLEDGMENTS

This work was partially supported by the CNPq, INCT-FCx, and by the US-AFOSR under the Grant No. FA9550-16-1-0280.

- <sup>1</sup>P. Ewald, *Ann. Phys.* **369**, 253 (1921).
- <sup>2</sup>J. Kolafa and J. W. Perram, *Mol. Simul.* **9**, 351 (1992).
- <sup>3</sup>U. Essmann, L. Perera, M. L. Berkowitz, T. Darden, H. Lee, and L. G. Pedersen, *J. Chem. Phys.* **103**, 8577 (1995).
- <sup>4</sup>T. Darden, D. York, and L. Pedersen, *J. Chem. Phys.* **98**, 10089 (1993).
- <sup>5</sup>D. Frenkel and B. Smit, *Understanding Molecular Simulation: From Algorithms to Applications* (Academic Press, San Diego, California, USA, 2002), pp. 291–306.
- <sup>6</sup>J. Lekner, *Phys. A* **176**, 485 (1991).
- <sup>7</sup>A. H. Widmann and D. B. Adolf, *Comput. Phys. Commun.* **107**, 167 (1997).
- <sup>8</sup>M. Mazars, *Mol. Phys.* **103**, 1241 (2005).
- <sup>9</sup>M. Deserno and C. Holm, *J. Chem. Phys.* **109**, 7678 (1998).
- <sup>10</sup>E. Spohr, *J. Chem. Phys.* **107**, 6342 (1997).
- <sup>11</sup>J. Hautman and M. L. Klein, *Mol. Phys.* **75**, 379 (1992).
- <sup>12</sup>M. Kawata and M. Mikami, *Chem. Phys. Lett.* **340**, 157 (2001).
- <sup>13</sup>A. Arnold, J. de Joannis, and C. Holm, *J. Chem. Phys.* **117**, 2496 (2002).
- <sup>14</sup>T. C. Yeh and M. L. Berkowitz, *J. Chem. Phys.* **111**, 3155 (1999).
- <sup>15</sup>A. P. dos Santos, M. Giroto, and Y. Levin, *J. Chem. Phys.* **144**, 144103 (2016).
- <sup>16</sup>J. F. Wishart, *Energy Environ. Sci.* **2**, 956 (2009).
- <sup>17</sup>Q. Li, Q. Tang, B. He, and P. Yang, *J. Power Sources* **264**, 83 (2014).
- <sup>18</sup>P. Simon and Y. Gogotsi, *Nat. Mater.* **7**, 845 (2008).
- <sup>19</sup>J. Chmiola, C. Largeot, P. L. Taberna, P. Simon, and Y. Gogotsi, *Science* **328**, 480 (2010).
- <sup>20</sup>D. T. Limmer, C. Merlet, M. Salanne, D. Chandler, P. A. Madden, R. van Roij, and B. Rotenberg, *Phys. Rev. Lett.* **111**, 106102 (2013).
- <sup>21</sup>J. I. Siepmann and M. Sprik, *J. Chem. Phys.* **102**, 511 (1995).
- <sup>22</sup>S. Reed, O. Lanning, and P. Madden, *J. Chem. Phys.* **126**, 084704 (2007).
- <sup>23</sup>C. Merlet, D. T. Limmer, M. Salanne, R. van Roij, P. A. Madden, D. Chandler, and B. Rotenberg, *J. Chem. Phys. C* **118**, 18291 (2014).
- <sup>24</sup>M. V. Fedorov and A. A. Kornishev, *Chem. Rev.* **114**, 2978 (2014).
- <sup>25</sup>M. Giroto, T. Colla, A. P. dos Santos, and Y. Levin, *J. Phys. Chem. B* **121**, 6408 (2017).
- <sup>26</sup>H. Tokuda, K. Hayamizu, K. Ishii, Md. A. B. H. Susan, and M. Watanabe, *J. Phys. Chem. B* **108**, 16593 (2005).

- <sup>27</sup>H. Tokuda, K. Hayamizu, K. Ishii, Md. A. B. H. Susan, and M. Watanabe, *J. Phys. Chem. B* **109**, 6103 (2005).
- <sup>28</sup>H. Tokuda, K. Ishii, Md. A. B. H. Susan, S. Tsuzuki, K. Hayamizu, and M. Watanabe, *J. Phys. Chem. B* **110**, 2833 (2006).
- <sup>29</sup>J. D. Jackson, *Classical Electrodynamics* (Wiley, 1999), pp. 140–141.
- <sup>30</sup>Y. Levin, *Europhys. Lett.* **76**, 163 (2006).
- <sup>31</sup>V. Kobelev, A. B. Kolomeisky, and M. E. Fisher, *J. Chem. Phys.* **116**, 7589 (2002).
- <sup>32</sup>M. C. Lonergan, J. W. Perram, M. A. Ratner, and D. F. Shriver, *J. Chem. Phys.* **98**, 4937 (1993).
- <sup>33</sup>H. Bartsch, O. Dannenmann, and M. Bier, *Phys. Rev. E* **91**, 042146 (2015).
- <sup>34</sup>*Ionic Liquids in Synthesis*, edited by P. Wasserscheid and T. Welton (Wiley-VCH, Weinheim, Germany, 2003).
- <sup>35</sup>M. N. Kobrak, *Green Chem.* **10**, 80 (2008).
- <sup>36</sup>M.-M. Huang, Y. Jiang, P. Sasisanker, G. W. Driver, and H. Weingartner, *J. Chem. Eng. Data* **56**, 1494 (2011).
- <sup>37</sup>J. MacDonald and C. Barlow, Jr., *J. Chem. Phys.* **36**, 3062 (1962).
- <sup>38</sup>Y. Levin, *Rep. Prog. Phys.* **65**, 1577 (2002).
- <sup>39</sup>A. A. Kornyshev, *J. Phys. Chem. B* **111**, 5545 (2007).
- <sup>40</sup>K. Breitsprecher, P. Košován, and C. Holm, *J. Phys.: Condens. Matter* **26**, 284108 (2014).
- <sup>41</sup>K. Breitsprecher, P. Košován, and C. Holm, *J. Phys.: Condens. Matter* **26**, 284114 (2014).